**Synthesis And Characterization Of Mixed Sulphadiazine - Ampicillin Complexes**

Bamigboye M.O\*1 and Isaac Y. A2

1. Department of Industrial Chemistry, P.M.B 1515, University of Ilorin, Nigeria

2. Department of Chemistry, P.M.B 1515, University of Ilorin, Nigeria

seymercyx@yahoo.com

**Abstract**: Three metal complexes of [Cu(Su)(Amp)Cl2], [Co(Su)(Amp)Cl2], and [Cd(Su)(Amp)Cl2] were synthesized using concentrating and mixing methods. The complexes formed were characterized by melting point, conductivity and infrared spectroscopy. An octahedral geometry was assigned for all the synthesized complexes. The ligands used were found to be bidentate. In sulphadiazine, coordination occurs through the nitrogen of the pyrimidine group and Sulphone group while in Ampicillin, coordination occurs through oxygen of the carbonyl group shown on the Infrared spectra. Investigation of the antimicrobial study was also carried out which indicates that the synthesized complexes exhibits higher activities than their parent ligands. The aim of the research work is to synthesized new alternative drugs which will be more effective than their parent ligands.

**[**Bamigboye M.O and Isaac Y. A. **Synthesis And Characterization Of Mixed Sulphadiazine - Ampicillin Complexes.** *Nat Sci* 2015;13(4):79-82]. (ISSN: 1545-0740). <http://www.sciencepub.net/nature>. 11

**Keywords:** Sulphadiazine, Ampicillin, Synthesis, Antimicrobial studies

1. **Introduction**

Previous reports have shown that efficacy of a therapeutic agent (such as sulphadiazine) increased upon coordination to transition metals, [Ajibola, 1990 and Akanji *et al*., 1993]. Sulphadiazine is an antimalarial drug and is used in treatment of malarial due to its esquizonticide effect, [Craig *et al*., 1982]. The short half-life of antibiotics (such as ampicillin), has proved to provide the benefits of therapy in rapid clearing bacterial disease that results in the release of harmful metals into blood, [Calvin *et al*., 1953]. Ampicillin is an antibiotic and it plays an important role in maintaining the level of bacterial causing disease in the body, [Hershko *et al*., 1994]. Key Pour *et al.,* 1986 reported the use of ascorbic acid as a possible antidote for iron overload in reaction of ferric iron with ascorbic acid. Bessman *et al* ., 1954 reported the use of calcium-EDTA as an effective remedy for lead poisoining. The non-toxic penicillamine is effectively used to remove excess copper in patients suffering from Wilsons disease, [Adrien *et al*.,1985].

Few reports regarding metal complexes of mixed sulphadiazine and ampicillin as possible antidotes for metal overload in the body appeared in literature. It is in line with this observation that we investigate the metal-chelating ability of mixed sulphadiazine and ampicillin ligands, and to report the antimicrobial activities for possibility regarding their usage as metal-chelator therapy.

1. **Experimental**

**Materials and Instrumentation:** All reagents and chemicals were of analytical grade and used as obtained from Aldrich. Sulphadiazine and Ampicillin were obtained as gift from Sam Pharmaceutical Ltd, Ilorin, Nigeria. Metal salts used include Copper chloride dehydrate [CuCl2.2H2O], Cobalt chloride hexahydrate [CoCl2.6H2O] and Cadmium (II) chloride [CdCl2.21/2H2O]. IR spectra of the samples in KBr pellets were obtained in the ranges of 4000-400 cm-1 using FTIR spectrometer. Metal Analyses were determined by atomic absorption spectroscopy with Perkin-Elmer Spectrometer, model 3110. UV-Vis spectra were obtained on Aquamate v4.60 spectrophotometer.

**Experimental Procedure**

**Synthesis of [Cu(Su)(Amp)Cl2] complex:** 2mmol of each of sulphadiazine and ampicillin were dissolved in 10ml of each of acetone and methanol in one conical flask and 1mmol of [CuCl2.2H2O] in 10ml of ethanol in another flask. The clear solution observed after mixture was concentrated at 50oC. Precipitate formed was allowed to cool, was filtered, washed and dried. The product was labelled for analysis/characterization.

MX2.nH2O Ligand → M(Ligand)2

**Synthesis of [Co(Su)(Amp)Cl2] complex:** 2mmol of each of sulphadiazine and ampicillin were dissolved in 10ml of each of acetone and methanol in separate conical flasks and 1mmol of [CoCl2.2H2O], in 10ml of ethanol in another flask. The clear solution observed after mixture was concentrated at 50oC. Precipitate formed was allowed to cool, filtered, washed and dried. The product was labelled for analysis/characterization.

**Synthesis of [Cd(Su)(Amp)Cl2] complex:** 2mmol of each of sulphadiazine and ampicillin were dissolved in 10ml of each of acetone and methanol in separate conical flasks and 1mmol of [CdCl2.2H2O], in 10ml of ethanol in another flask. After mixing a clear solution of the mixed ligand, and metal salt, was left to stand for 24hrs. Precipitate formed was allowed to cool, filtered, washed and dried. The product was labelled for analysis/characterization.

**3. Antimicrobial Test:** The inhibitory action of all the complexes was tested against *Staphylococcus aureus, Escherichia coli,* and *Klebsiella pneumonia.*

**4. Results And Discussion**

The Cu (II), Co (II) and Cd (II) complexes of sulphadiazine-ampicillin ligands were synthesized by reaction of metal salts with mixed sulphadiazine and ampicillin. The complexes were characterized by AAS, Conductivity, TLC, infrared, UV-Vis spectroscopy and antimicrobial studies was investigated. The complexes are generally soluble in methanol, ethanol, acetone and DMSO but insoluble in non-polar organic solvents. The complexes are non-hygroscopic solids with melting points higher than the parent drug. The physical properties of the various complexes are collected in Table 1. The molar conductance values measured in DMSO solution (10-3 M) for these complexes are 7.31 x 10-6, 6.58 x 10-6, and 6.46 x 10-6 for the Cu, Co and Cd complexes respectively (Table 1). Similarly, from the results, the complexes are non-electrolytes. Determination of stoichiometric ratio using job’s method [Hubel., 2000] suggests a 1:2 metal to ligand stoichiometry for the complexes. The proposed structures of the mixed complexes are shown below.

**TABLE 1:** **Analytical Data** **of** **Sulphadiazine and Ampicillin metal complexes**

|  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- |
| Ligand/mixture | Colour | Melting pointoC | ConductivityΩ-1 mol-1 cm2 | % Metal | TLC Rf |
| Sulphadiazine (ligand) | White | 230 | 6.4 x 10-6 |  | 0.34 |
| Ampicillin (ligand) | White | 237-238 | 2.8 x 10-6 |  | 0.37 |
| [Cu(Su)(Amp)Cl2] | Blue | 247 | 6.46 x 10-6 | 10.30 | 0.35 |
| [Co(Su)(Amp)Cl2] | Pink | 270 | 6.58 x 10-6 | 15.60 | 0.34 |
| [Cd(Su)(Amp)Cl2] | White | 251 | 7.31 x 10-6 | 17.70 | 0.36 |

**UV-Visible Spectroscopy Results**

**Table 2: UV-Visible spectra of Sulphadiazine and Ampicillin metal complexes**

|  |  |  |  |
| --- | --- | --- | --- |
| Ligand/mixture | *Wavelength(nm)* | *Energies(cm-1)* | *Assignment* |
| Sulphadiazine (ligand) | 268325244 | 37313.430769.240983.6 | π -π\*LCTn-π\* |
| Ampicillin (ligand) | 229325 | 43668.130769.5 | n-π\*n-π\*n-π\* |
| [Cu(Su)(Amp)Cl2] | 821 | 12180 | 2Eg - 2T2g |
| [Co(Su)(Amp)Cl2] | 367585850 | 272471223911764 | MLCT4T1g -4A2g4T1g - 4T2g |
| [Cd(Su)(Amp)Cl2] | 193256 | 51813.538940.8 | n-π\*n-π\* |

The UV-Vis. Spectra of sulphadiazine in methanol present three absorption bands maxima at 268 and 325 which is assigned to π -π\* transition, 325 is assigned to LCT and 244 is assigned to n-π\* transitions. The UV-Vis spectra of ampicillin in acetone present two absorption band maxima at 229 and 325 which is assigned to n-π\* transitions respectively. Copper(II) complex exhibits one band centered at 821 which may be tentatively assigned to 2Eg - 2T2g transitions respectively. The cobalt (II) complex exhibits 3 bands at 367, 585 and 850 nm. These bands are assigned to MLCT, 4T1g - 4A2g and 4T1g -4T2g transitions respectively. Cadium (II) exhibits two absorption bands centered at 193 and 256 which may be assigned to n-π\* transition respectively. All these characteristic bands observed in the UV-Vis. spectra confirm octahedral configuration for Co(II), Cd(II) and Cu(II) complexes respectively.

**Infrared Spectroscopy Results**

**Table 3: Infrared spectra result for mixed Su-Amp metal complexes**

|  |  |  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- | --- | --- |
| **Ligand/Complexes** | **N-H** | **(SO2)** | **(C=O)** | **(C-O)** | **(C-N)** | **C=N** | **C-S** |
| Sulphadiazine (ligand) | 3424.02s | 1325.76w,b | 1262.31s | 1187.79s | 1157.77b | 1681.04s | - |
| Ampicillin (ligand) | 3525.80 m,b | - | 1396.79s | 874.19w,b | 1373.60s | 1684.29s | - |
| [Cu(Su)(Amp)Cl2] | 3538.68s,s | 1384.82s,b | 1268.55s | 1126.52m,b | 1368.67s | 1629.94s | 697.52 |
| [Co(Su)(Amp)Cl2] | 3413.14s | 1325.93m,b | 1156.61s | - | 1387.76s | 1652.17w | 687.45b |
| [Cd(Su)(Amp)Cl2] | 3403.53s | 1326.55s | 1557.77s | 1157.77s | 1387.56s | 1641.04s | 689.10b |

Sulphadiazine and Ampicillin can provide two donor atoms to the central atom due to steric constraints. The coordination of sulphadiazine occurs through the nitrogen of the pyrimidine group and Sulphone group while in Ampicillin, coordination occurs through oxygen of the carbonyl group shown on the Infrared spectra. The values of IR spectra of both ligands and their complexes have been assigned mainly for those specific frequencies directly involved in complex formation, (Table 3). The vibrational band at 1681.04cm-1 C=N Sulphadiazine is shifted to lower wave number 1629.94 cm-1 for Cu (II) complex, 1652.17 cm-1 for Co (II) complex, and 1641.04cm-1 for Cd(II) complex, thus indicating coordination through C=N at 3424.02cm-1 N-H in sulphadiazine shifted to higher wave number (3538.68cm-1) in [Cu(Su)(Amp)Cl2] complex; and finally shifted to lower wave numbers (3413.14 cm-1 and 3403.53cm-1) in [Co(Su)(Amp)Cl2] and [Cd(Su)(Amp)Cl2] complexes; this is probably due to coordination of metal via Nitrogen of N-H group of sulphadiazine in the complexes.

Special attention is also paid to the bands at 1325.76cm -1 assigned to sulphone group ν(S=O) in sulphadiazine which shifted to higher wave number (1384.82 cm-1) for Cu(II) complex while the bands at (1325.93 cm -1 and 1326.55 cm -1) assigned to (S=O) shows slight regularity as well for Co(II) and Cd(II) complexes, thus indicating that this group may or may not be involved in coordination with the Cu(II), Co(II) and Cd(II) complexes. Spectra changes observed from 1373.60cm-1 in ampicillin to (1368.67 cm-1, 1387.76 cm-1 and 1387.56 cm-1) corresponding to ν(C-N) of the nitrile group in Cu(II), Co(II) and Cd(II) is so insignicant, indicating that this group is not involved in coordination of all the complexes. Similarly, the insignificant spectra changes observed at (697.52cm -1, 687.45cm -1 and 689.10cm -1) corresponding to ν(C-S) for all the complexes and for which none appeared in ampicillin, suggests there is no possibility of coordination at this site to the Cu(II), Co(II) and Cd(II) ions. However, the band at 1396.79cm-1 corresponding to ν(C=O) in ampicillin shifted to higher wave numbers (1268.55cm -1, 1156.61 cm -1 and 1557.77cm -1) in all the complexes thus, indicating that this group is possibly involved in coordination of all complexes.

From the result in Table 4 above, it is evident that the overall zone of inhibition against bacterial species is highest in Co(II) complex, followed by Cu(II) complex while Cd(II) complex appears to have the least zone of inhibition. It is also evident that nearly all the metal complexes posses significant zone of inhibition acting against bacterial species than their parent ligands thus, giving the metal-chelators potentiality as antidotes for metal–overload.

**Table 4: Antimicrobial Activity of Sulphadiazine and Ampicillin metal complexes**

|  |  |  |  |
| --- | --- | --- | --- |
| **Ligand/Complex** | **ʹEscherichia coli (nm)(-ve)** | **Staphylococcus (nm) (-ve)** | **Klebsiella (nm) (-ve)** |
| Concentration (ppm) | 20 | 40 | 50 | 100 | 20 | 40 | 50 | 100 | 20 | 40 | 50 | 100 |
| Sulphadiazine (ligand) | 0 | 0 | 0 | 7.8 | 0 | 9.7 | 0 | 3.50 | 0 | 4.5 | 0 | 0 |
| Ampicillin (ligand) | 0 | 3.4 | 0 | 0 | 2.5 | 0 | 0 | 0 | 4.5 | 0 | 0 | 0 |
| [Cu(Su)(Amp)Cl2] | 26.8 | 22.2 | 13.5 | 37.0 | 0 | 0 | 0 | 0 | 15.5 | 17.09 | 7.8 | 17.09 |
| [Co(Su)(Amp)Cl2] | 23.4 | 25.6 | 14.4 | 38.5 | 0 | 0 | 0 | 0 | 16.5 | 15.7 | 7.2 | 15.3 |
| [Cd(Su)(Amp)Cl2] | 0 | 0 | 34.9 | 35.6 | 0 | 0 | 0 | 3.1 | 15.7 | 16.98 | 8.9 | 16.9 |



[M(Su)(Amp)Cl2]



[M(Su)(Amp)Cl2]

**Proposed Structure**

The results obtained from spectroscopic analysis for this study revealed that coordination of sulphadiazine, occurs through the nitrogen of the pyrimidine group and Sulphone group while in Ampicillin, coordination occurs through oxygen of the carbonyl group for all the metal complexes of [Cu(Su)(Amp)Cl2], [Co(Su)(Amp)Cl2], and [Cd(Su)(Amp)Cl2] synthesized to give an octahedral geometry, with ligand acting as bidentate. This mode of coordination correspond to those in literature [Keypour et al., 1986 and Bessman et al 1954].

Evidence of formation of complex was demonstrated by examining the IR and UV spectra carefully.

**5. Conclusion**

The metal complexes of [Cu(Su)(Amp)Cl2], [Co(Su)(Amp)Cl2], and [Cd(Su)(Amp)Cl2] have been synthesized by carefully mixing two ligands (sulphadiazine and ampicilling) in suitable solvents and adding suitably dissolved inorganic salts of Cu(II), Co(II) and Cd(II) into solution of ligands respectively. The structure of the product formed was characterized by AAS, conductivity, TLC, infrared and UV-Vis spectroscopy. In all the complexes formed, sulphadiazine coordinated through the nitrogen of the pyrimidine group and Sulphone group while Ampicillin coordinated through oxygen of the carbonyl group. The mixed ligands being bidendate is assumed to coordinate with the three metal ions (copper, cobalt and cadmium) thus, forming an octahedral geometry. Two donors atom (Nitrogen atom of pyrimidine group and sulphur atom of sulphone group in sulphadiazine and oxygen atom of carbonyl group in ampicillin) are involved in coordination of sulphadiazine-ampicillin with the three metal ions. In summary, it is evident that the overall antimicrobial activity is high for the [Cu(Su)(Amp)Cl2], [Co(Su)(Amp)Cl2], and [Cd(Su)(Amp)Cl2] complexes thus giving the three complexes potentiality as antidotes for metal–overload/poisoning.

**References**

1. Ajibola A.O (1990). Essential of Medicinal Chemistry; 2nd edition; Sharson Jersey, pp. 28-446.
2. Akanji M.A, Olagoke O.A, Oloyede O.B (1993). Effect of chronic consumption of metabisulphite on the integrity of the kidney cellular system. Toxicol. 81: 173-179.
3. Craig C.R, Stitzel R.E (1982). Modern Pharmacology. 5th edition. Little Brown and Co., Boston. pp. 653-655.
4. Calvin M.and A.E. Martell, Chemistry of metal chelate compounds, Prentice Hall, Inc., New York, 2nd Ed. (1953) p. 134.
5. Hershko, C. Link,G. Tzahor M. and Pinson, A. Metal-based drugs. 1, 2-3,83(1994).
6. Keypour, H. Silver, M.T. Wilson, M.T. Hamed, (1986). Inorg. Chim. Acta, 125, 97.
7. Bessman, S.P. Dublin, M, and Leikin , S. (1954) Paediatrics., 14, 201.
8. Adrien, A. (1985). *Selective Toxicity*, 7th Ed. Chapman and Hall, New York, 432-468.
9. Hubel, R. Jelinek T. and Beck, W. (2000) *Zeitschrift fuer Naturforchung*, , 55b, 821.

3/16/2015