

The Best Periodic Table Of The Elements

Jiang Chun-xuan

P. O. Box 142-206, Beijing 100854, P. R. China

j136211668701@163.com

Abstract: Using the stable number theory[3] we make the best periodic table of the elements and obtain the best electron configurations of the elements.

[Jiang Chun-xuan. **The Best Periodic Table Of The Elements**. *Rep Opinion* 2025;17(11):1-9]. ISSN 1553-9873 (print); ISSN 2375-7205 (online). <http://www.sciencepub.net/report>. 01. doi:[10.7537/marsroj171125.01](https://doi.org/10.7537/marsroj171125.01)

Keywords: Best; Periodic Table; Elements; electron configuration

In studying the stability of the many-body problem we suggest two principles [1-9].

(1) The prime number principle. A prime number is irreducible in the integers, it seems therefore natural to associate it with the most stable subsystem. We prove that 1, 3, 5, 7, 11, 23, 47 are the most stable primes.

(2) The symmetric principle. The most stable configuration of two prime numbers is then stable symmetric system in nature. We prove that 2, 4, 6, 10, 14, 22, 46, 94 are the most stable even numbers. The stability can be defined as long life and existence in nature, and instability as short life or non-existence in nature.

In this paper by using the prime number principle and the symmetric principle we make the new electron configurations of the elements. Total quantum number n and orbital quantum number l determine the new electron configurations of the elements

Electron shells:	$n=1$	2	3	4	5	6...
	<i>K</i>	<i>L</i>	<i>M</i>	<i>N</i>	<i>O</i>	<i>P</i> ...

Electron subshells:	$2(2l+1)=2$	6	10	14	18	22...
	<i>s</i>	<i>p</i>	<i>d</i>	<i>f</i>	<i>g</i>	<i>h</i> ...

An atomic subshell that contains its full quota of electrons is said to be closed. A closed s subshell ($l=0$) holds two electrons, a closed p subshell ($l=1$) six electrons, a closed d subshell ($l=2$) ten electrons, a closed f subshell ($l=3$) fourteen electrons, these subshells are the most stable, a closed g subshell ($l=4$) eighteen electrons is the most unstable. Using the symmetric principle it has been proved the $2(2l+1)=2, 6, 10$ and 14 are stable and $2(2l+1)=18$ is unstable. The s, p, d , and f subshells are stable and the g subshell is unstable.

Table 1 shows the best electron configurations of the elements. From 1 to 92 of the atomic numbers every subshell is stable. It has been proved that the last stable element that occurs naturally is uranium with an atomic number of 92 and there are only 92 stable elements in nature. Since $5g$ subshell is unstable, the elements 93-110 are unstable. Since $5g$ is unstable, $6s, 6p, 6d, 6f, 6g$ and $6h$ subshells are unstable. Therefore the elements 111-182 are unstable.

Many of the chemical and physical properties of the elements are related to the number of electrons in the outermost shell, the electrons that are valence electrons in these atoms. In table 1 there are correct valence electron configurations of the elements. In Mendeleev periodic table the elements (1-18, 29-36 and 46) have correct valence electron configurations and the elements (19-28, 37-45 and 47-92) have wrong valence electron configurations. From table 1 we make new periodic table of the elements.

Conclusion. In table 1 s, p, d and f are stable subshells. Therefore the elements 1 to 92 are stable. In table 1 $5g$ is the unstable subshell. Therefore the elements 93 to 110 are unstable. Using the table 1 and 2 chemists study the chemical properties of the elements 1 to 92 and discover many of the new chemical compounds. Pythagoras believes that all things are numbers. But Jiang believes that all things are stable numbers [1-9].

这是人类第一次用数论证明并给出新元素电子结构和正确价电子结构, 这是化学基础. 化学家研究新元素电子结构将会发现很多新化合物. 门捷列夫元素周期表是错的, 元素 1-18, 用实验证明价电子是正确的, 估算元素 29-36 和 46 价电子是正确的, 但其它元素价电子是不正确的, 元素没有周期性. 目前化学研究是猜想没有理论指导. 本文是蒋春暄 1981 年提出"素数原理和对称原理" 两个原理一个应用. 化学是计算问题. 只能用正确价电子才能找到新化合物. 本文给化学和生物化学提供一个理论基础. 首先理论计算而后在实验室可找到这种化合物和生物结构. 人类研究结果自然界只有 92 种元素, 所有科学家对这个问题一筹莫展, 本文用非常简单方法就证明了自然界只有 92 种元素, 其它元素都是人造的. 所以这种证明是正确的. 化学理论要重新改写. 本文已在国外上网, 这应该是"十八大" 最大成果, 我最近才知 **valence electrons** 的重要性, 所以重写本文. 元素电子排列应该是连续的, 门捷列夫周期表中电子排列不是连续的, 对元素 >18 核外电子排列按周期性排列是跳跃式排列, 中间留下断层, 他们也知道 $5g$ 和 $6g$ 是不稳定的, 排列元素电子不用 $5g$ 和 $6g$, 但元素电子排列利用不稳定 $6s, 6p, 6d, 7s, \dots$ 这是错误的. 这个问题延续到今天仍是这样. 房子只能一层一层的建, 保证每层都稳是才行. 第五层不稳定, 那末第六层也是不稳定的, 这道理大家都明白. 我们证明 $5g$ 是不稳定的, 所以元素 93-110 是不稳定的, 所以 $6s, 6p, 6d, 6f, 6g, 6h$ 是不稳定的, 以后元素更不稳定的. 我们证明自然界只有 92 种稳定元素是绝对正确的.

Table 1. Jiang Electron Configuration of the Elements

Z	Sym	K		L			M			N				O				
		1s	2s	2p	3s	3p	3d	4s	4p	4d	4f	5s	5p	5d	5f	5g		
1	H	1																
2	He	2																

3	Li	2	1							
4	Be	2	2							
5	B	2	2	1						
6	C	2	2	2						
7	N	2	2	3						
8	O	2	2	4						
9	F	2	2	5						
1	Ne	2	2	6						
0	Na	2	2	6	1					
1	Mg	2	2	6	2					
1	Al	2	2	6	2	1				
1	Si	2	2	6	2	2				
2	P	2	2	6	2	3				
1	S	2	2	6	2	4				
3	Cl	2	2	6	2	5				
1	Ar	2	2	6	2	6				
4	K	2	2	6	2	6	1			
1	Ca	2	2	6	2	6	2			
5	Sc	2	2	6	2	6	3			
1	Ti	2	2	6	2	6	4			
6	V	2	2	6	2	6	5			
1	Cr	2	2	6	2	6	6			
7	Mn	2	2	6	2	6	7			
1	Fe	2	2	6	2	6	8			
8	Co	2	2	6	2	6	9			
1	Ni	2	2	6	2	6	10			
9	Cu	2	2	6	2	6	10	1		
2	Zn	2	2	6	2	6	10	2		
0	Ga	2	2	6	2	6	10	2	1	
2	Ge	2	2	6	2	6	10	2	2	
1	As	2	2	6	2	6	10	2	3	
2	Se	2	2	6	2	6	10	2	4	
2	Br	2	2	6	2	6	10	2	5	
2	Kr	2	2	6	2	6	10	2	6	
3	Rb	2	2	6	2	6	10	2	6	1
2	Sr	2	2	6	2	6	10	2	6	2
4	Y	2	2	6	2	6	10	2	6	3
2	Zr	2	2	6	2	6	10	2	6	4
5	Nb	2	2	6	2	6	10	2	6	5
2	Mo	2	2	6	2	6	10	2	6	6
6	Tc	2	2	6	2	6	10	2	6	7

2	Ru	2	2	6	2	6	10	2	6	8	
7	Rh	2	2	6	2	6	10	2	6	9	
2	Pd	2	2	6	2	6	10	2	6	10	
8											
2											
9											
3											
0											
3											
1											
3											
2											
3											
3											
3											
4											
3											
5											
3											
6											
3											
7											
3											
8											
3											
9											
4											
0											
4											
1											
4											
2											
4											
3											
4											
4											
4											
5											
4											
6											

Table 1. (Continued)

Z	Sym	K			L			M			N				O				
		1s	2s	2p	3s	3p	3d	4s	4p	4d	4f	5s	5p	5d	5f	5g			
47	Ag	2	2	6	2	6	10	2	6	10	1								
48	Cd	2	2	6	2	6	10	2	6	10	2								
49	In	2	2	6	2	6	10	2	6	10	3								
50	Sn	2	2	6	2	6	10	2	6	10	4								
51	Sb	2	2	6	2	6	10	2	6	10	5								
52	Te	2	2	6	2	6	10	2	6	10	6								
53	I	2	2	6	2	6	10	2	6	10	7								
54	Xe	2	2	6	2	6	10	2	6	10	8								
55	Cs	2	2	6	2	6	10	2	6	10	9								
56	Ba	2	2	6	2	6	10	2	6	10	10								
57	La	2	2	6	2	6	10	2	6	10	11								
58	Ce	2	2	6	2	6	10	2	6	10	12								
59	Pr	2	2	6	2	6	10	2	6	10	13								
60	Nd	2	2	6	2	6	10	2	6	10	14								
61	Pm	2	2	6	2	6	10	2	6	10	14	1							
62	Sm	2	2	6	2	6	10	2	6	10	14	2							
63	Eu	2	2	6	2	6	10	2	6	10	14	2	1						
64	Gd	2	2	6	2	6	10	2	6	10	14	2	2						
65	Tb	2	2	6	2	6	10	2	6	10	14	2	3						
66	Dy	2	2	6	2	6	10	2	6	10	14	2	4						
67	Ho	2	2	6	2	6	10	2	6	10	14	2	5						
68	Er	2	2	6	2	6	10	2	6	10	14	2	6						
69	Tm	2	2	6	2	6	10	2	6	10	14	2	6	1					
70	Yb	2	2	6	2	6	10	2	6	10	14	2	6	2					
71	Lu	2	2	6	2	6	10	2	6	10	14	2	6	3					
72	Hf	2	2	6	2	6	10	2	6	10	14	2	6	4					
73	Ta	2	2	6	2	6	10	2	6	10	14	2	6	5					
74	W	2	2	6	2	6	10	2	6	10	14	2	6	6					
75	Re	2	2	6	2	6	10	2	6	10	14	2	6	7					
76	Os	2	2	6	2	6	10	2	6	10	14	2	6	8					
77	Ir	2	2	6	2	6	10	2	6	10	14	2	6	9					
78	Pt	2	2	6	2	6	10	2	6	10	14	2	6	10					

79	Au	2	2	6	2	6	10	2	6	10	14	2	6	10	1
80	Hg	2	2	6	2	6	10	2	6	10	14	2	6	10	2
81	Tl	2	2	6	2	6	10	2	6	10	14	2	6	10	3
82	Pb	2	2	6	2	6	10	2	6	10	14	2	6	10	4
83	Bi	2	2	6	2	6	10	2	6	10	14	2	6	10	5
84	Po	2	2	6	2	6	10	2	6	10	14	2	6	10	6
85	At	2	2	6	2	6	10	2	6	10	14	2	6	10	7
86	Rn	2	2	6	2	6	10	2	6	10	14	2	6	10	8
87	Fr	2	2	6	2	6	10	2	6	10	14	2	6	10	9
88	Ra	2	2	6	2	6	10	2	6	10	14	2	6	10	1
89	Ac	2	2	6	2	6	10	2	6	10	14	2	6	10	0
90	Th	2	2	6	2	6	10	2	6	10	14	2	6	10	1
91	Pa	2	2	6	2	6	10	2	6	10	14	2	6	10	1
92	U	2	2	6	2	6	10	2	6	10	14	2	6	10	1
															2
															1
															3
															1
															4

Table 1.(Continued)

Z	Sym	K			L			M			N				O				
		1s	2s	2p	3s	3p	3d	4s	4p	4d	4f	5s	5p	5d	5f	5g			
93	Np	2	2	6	2	6	10	2	6	10	14	2	6	10	1	1			
94	Pu	2	2	6	2	6	10	2	6	10	14	2	6	10	4	2			
95	Am	2	2	6	2	6	10	2	6	10	14	2	6	10	1	3			
96	Cm	2	2	6	2	6	10	2	6	10	14	2	6	10	4	4			
97	Bk	2	2	6	2	6	10	2	6	10	14	2	6	10	1	5			
98	Cf	2	2	6	2	6	10	2	6	10	14	2	6	10	4	6			
99	Es	2	2	6	2	6	10	2	6	10	14	2	6	10	1	7			
100	Fm	2	2	6	2	6	10	2	6	10	14	2	6	10	4	8			
101	Md	2	2	6	2	6	10	2	6	10	14	2	6	10	1	9			
102	No	2	2	6	2	6	10	2	6	10	14	2	6	10	4	1			
103	Lr	2	2	6	2	6	10	2	6	10	14	2	6	10	1	0			
104	Rf	2	2	6	2	6	10	2	6	10	14	2	6	10	4	1			
105	Db	2	2	6	2	6	10	2	6	10	14	2	6	10	1	1			
106	Sg	2	2	6	2	6	10	2	6	10	14	2	6	10	4	1			
107	Bh	2	2	6	2	6	10	2	6	10	14	2	6	10	1	2			
108	Hs	2	2	6	2	6	10	2	6	10	14	2	6	10	4	1			
109	Mt	2	2	6	2	6	10	2	6	10	14	2	6	10	1	3			

110	Ds	2	2 6	2 6 10	2 6 10 14	2 6 10 4 1	1 4
							4 1
							1 5
							4 1
							1 6
							4 1
							1 7
							4 1
							1 8
							4
							1
							4
							1
							4
							1
							4

Table 2. The Best periodic table of elements.

Atomic Orbitals	Outermost Subshell electrons	1. Period	2. Period	3. Period	4. Period	5. Period
s	1	1 H	3 Li	11 Na	29 Cu	61 Pm
	2	2 He	4 Be	12 Mg	30 Zn	62 Sm
p	1	Stable elements	5 B	13 Al	31 Ga	63 Eu
	2		6 C	14 Si	32 Ge	64 Gd
	3		7 N	15 P	33 As	65 Tb
	4		8 O	16 S	34 Se	66 Dy
	5		9 F	17 Cl	35 Br	67 Ho
	6		10 Ne	18 Ar	36 Kr	68 Er
d	1	Stable elements		19 K	37 Rb	69 Tm
	2		20 Ca	38 Sr	70 Yb	
	3		21 Sc	39 Y	71 Lu	
	4		22 Ti	40 Zr	72 Hf	
	5		23 V	41 Nb	73 Ta	
	6		24 Cr	42 Mo	74 W	
	7		25 Mn	43 Tc	75 Re	
	8		26 Fe	44 Ru	76 Os	

	9 10		27 Co 28 Ni	45 Rh 46 Pd	77 Ir 78 Pt
f	1 2 3 4 5 6 7 8 9 10 11 12 13 14	Stable elements		47 Ag 48 Cd 49 In 50 Sn 51 Sb 52 Te 53 I 54 Xe 55 Cs 56 Ba 57 La 58 Ce 59 Pr 60 Nd	79 Au 80 Hg 81 Tl 82 Pb 83 Bi 84 Po 85 At 86 Rn 87 Fr 88 Ra 89 Ac 90 Th 91 Pa 92 U
g	1 2 3 4 5 6 7 8 9 10 11 12 13 14 15 16 17 18	Unstable elements			93 Np 94 Pu 95 Am 96 Cm 97 Bk 98 Cf 99 Es 100 Fm 101 Md 102 No 103 Lr 104 Rf 105 Db 106 Sg 107 Bh 108 Hs 109 Mt 110 Ds

References

- [1] Jiang, Chun-xuan. A new theory for many-body problem stabilities. (Chinese) Qian Kexue 1, 38-48 (1981).
- [2] Jiang, Chun-xuan. On the symmetries and the stabilities of $4n + 2$ electron configurations of the elements. Phys. Lett. 73A, 385-386(1979).
- [3] Jiang, Chun-xuan. The application of stable groups to biological structures. Acta Math. Sci. 5, 243-260(1985).
- [4] Jiang, Chun-xuan. The prime principle in biology (Chinese), J. Biomath, 1, 123-125(1986).
- [5] Jiang, Chun-xuan. A mathematical model for particle classification. Hadronic J. Supp. 2, 514-522(1986).
- [6] Jiang, Chun-xuan. On the limit for the periodic table of the elements. Apeiron Vol. 5 Nr. 1-2, 21-24(1998).
- [7] Jiang, Chun-xuan. Foundations of Santilli's isonumber theory Part 1. Algebras, Groups and Geometries. 15, 351-393 (1998).
- [8] Jiang, Chun-xuan. Foundations of Santilli Isonumber Theory with applications to new cryptograms, Fermat's theorem and Goldbach's Conjecture.pp.85-88. Inter, Acad, Press. 2002. MR2004c:11001. <http://www.i-b-r.org/docs/jiang.pdf>
- [9] Jiang, Chun-xuan. The prime principle and the symmetric principle in clusters and nanostructures. <http://vixra.org/pdf/1004.0043v1.pdf>.

9/11/2025